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4. (ii) Mole in 12.046×10^{22} atoms of $\text{CaO} = \frac{12.046 \times 10^{23}}{6.022 \times 10^{23}}$

Mass of $\text{CaO} = 40 + 16 = 56 \text{ g/mole}$
 $\Rightarrow 2 \text{ mole of CaO} = 2 \times 56 \text{ g} = 112 \text{ g}$

5. (i) Mole of Cu in 12.046×10^{22} = $\frac{12.046 \times 10^{22}}{6.022 \times 10^{23}} = \frac{1.2046 \times 10^{23}}{6.022 \times 10^{23}} = \frac{1}{5} = 0.2 \text{ mole}$

(ii) Moles in 3.011×10^{23} molecules of substance = $\frac{3.011 \times 10^{23}}{6.022 \times 10^{23}} = \frac{1}{2} = 0.5 \text{ mole}$

(iii) Moles in 24.092×10^{23} molecules of $\text{H}_2\text{O} = \frac{24.092 \times 10^{23}}{6.022 \times 10^{23}} = 4 \text{ moles}$



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